Q1) An oxygen atom (O) contains 10 neutrons and 8 electrons.

1. Find the number of protons of oxygen atom.
2. Deduce its atomic number.
3. Find the mass number (A) of oxygen atom.
4. Some expected representation for oxygen atoms
   
   A: $^{16}_8O$  B: $^{18}_8O$  and C: $^{18}_{10}O$

Which of these three compounds are isotopes for oxygen atoms, (A&B) OR (B &C)? Justify?

Q2) 1. write the electron configuration for each of the following atoms:

11 Na
14 Si
6 C
2 He
17 Cl
2. specify the period, column, group & period for each of the above atoms (11 Na, 14 Si, 6 C, 2 He, 17 Cl, 19 F and 20 Ca).